For Supervisor's use only

90173





Level 1 Chemistry, 2003

90173 Describe the chemistry of selected non-metals and their compounds

Credits: Five 9.30 am Monday 24 November 2003

Check that the National Student Number (NSN) on your admission slip is the same as the number at the top of this page.

You should answer ALL the questions in this booklet.

If you need more space for any answer, use the pages provided at the back of this booklet and clearly number the question.

A metal activity series, solubility rules, a table of ions and a periodic table are provided in the Resource Booklet in your Level 1 Chemistry package.

Check that this booklet has pages 2–11 in the correct order and that none of these pages is blank.

YOU MUST HAND THIS BOOKLET TO THE SUPERVISOR AT THE END OF THE EXAMINATION.

Achievement Criteria	For Assessor's use only	
Achievement	Achievement with Merit	Achievement with Excellence
Describe the properties and preparation of, and state the importance of, selected nonmetals and their compounds.	Explain the properties, preparation and importance of selected non-metals and their compounds.	Analyse the properties and preparation of selected non-metals and their compounds, and evaluate their importance.
Overa	all Level of Performance	

You are advised to spend 50 minutes answering the questions in this booklet.

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QUESTION ONE: Sulfur and its Compounds

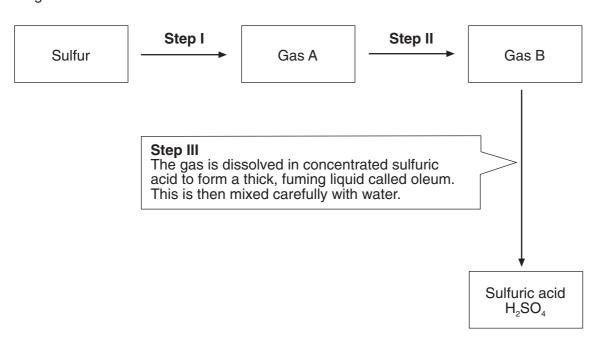
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Part	A: Sı	ılfur
(a)	Sulfu	ur is a by-product when natural gas is purified.
	Nam	e the compound in natural gas that is the source of this sulfur.
(b)	Nam	e and describe TWO allotropes of sulfur.
	(1)	Name:
		Description:
	(2)	Name:
		Description:

(a)

Name or give the formula of:

In New Zealand, sulfuric acid is made from sulfur by the contact process, summarised in the diagram below.



	(i)	Gas A:
	(ii)	Gas B:
(b)	(i)	Write an equation to describe the reaction occuring at Step I of the process.
	(ii)	Describe the conditions required for Step II and explain why these conditions are necessary.

(c) Containers and tankers of sulfuric acid display the warning sign at right.



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	What property of sulfuric acid makes this warning sign necessary?
d)	Explain why concentrated sulfuric acid can be stored in steel tanks but dilute sulfuric acid cannot.
Dort	C. Supembeenhete
Sulfu	C: Superphosphate ric acid is used to convert rock phosphate (calcium phosphate) into superphosphate, which is y used as a fertiliser on New Zealand farms.
Expla ertilis	ain why rock phosphate has to be converted to superphosphate before it is useful as a ser.

Cu + $HNO_3 \rightarrow$

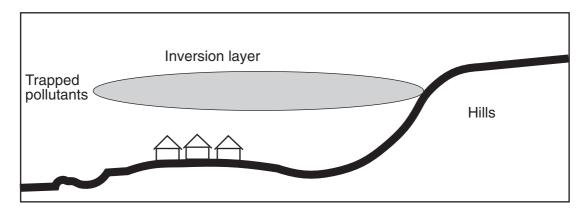
Part A: Nitrogen dioxide

Nitrogen dioxide gas is prepared in the laboratory by adding copper metal to concentrated nitric acid.

(1)	
(2)	
(=)	
(0)	
(3)	
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Part B: Nelson's pollution problems

Nelson is a small city surrounded by hills. It has a stable temperature inversion that traps ground-level air and allows air pollutant levels to build up.



Monitoring equipment measures the air pollution levels in Nelson. **Figure 1** below shows the readings taken between 28 July 2002 and 11 August 2002. The guideline maximum is 50 particles of solids per cubic metre of air.

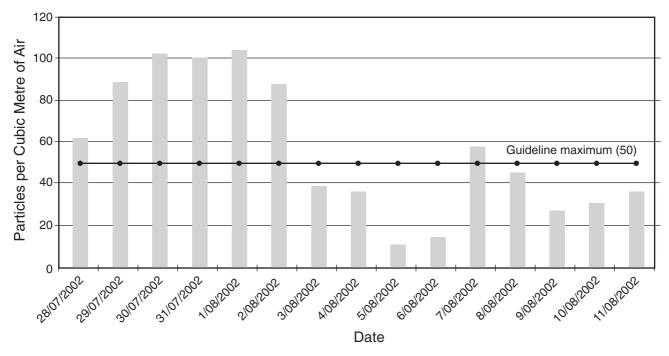


Figure 1

Pollutants from petrol and diesel engines are shown in **Figure 2** below. They are shown in grams of pollutant produced per kilogram of fuel.

Fuel	Carbon monoxide	Hydrocarbons	Nitrogen oxides	Sulfur dioxide	Particles of soot
Petrol	236	25	29	0.9	0.6
Diesel	10	17	59	3.8	18.0

Figure 2

Desc	ribe TWO effects of the presence of sulfur dioxide on people or the natural environment.
1)	
2)	
Discu	uss the relative contribution of petrol and diesel engines to photochemical smog .
	uss the effects of photochemical smog on both people and the environment, and the
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QUESTION THREE: Chlorine and its Compounds

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Chlorine gas can be made in the laboratory by the reaction of concentrated hydrochloric acid with sodium hypochlorite, NaOCI. **Figure 3** below shows how the preparation of chlorine can be carried out.

NaOCI + 2HCI \rightarrow NaCI + H₂O + CI₂

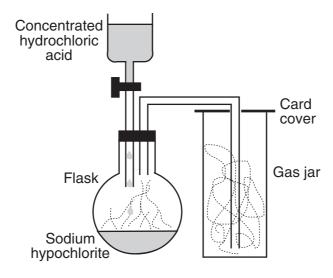


Figure 3

- (a) What colour is chlorine gas?
- (b) Another product of this reaction is NaCl. What is the chemical name of NaCl?
- (c) Explain why chlorine gas can be collected in the way shown in the diagram above rather than by water displacement.

(d) When NaCl is reacted with concentrated sulfuric acid, hydrogen chloride gas is formed.

Write a balanced chemical equation for this reaction.

QUESTION FOUR: Carbon Dioxide

Part A: Mineral water

A brand of bottled water is described as naturally carbonated, natural mineral water. This means that the fizz comes from carbon dioxide gas that is produced naturally underground rather than chemically manufactured gas added from a cylinder.



(a)	Write a balanced equation to show the reaction of carbon dioxide with water.
(b)	What effect would this solution have on litmus paper?
	B: A fire extinguisher
	d carbon dioxide is stored in fire extinguishers. When the pressure is released, it instantly erts to gas and some solid.
	the physical and chemical properties of carbon dioxide to its suitability for use in fire guishers.

Extra paper for continuation of answers if required. Clearly number the question.

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Question Number	

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